

Name _____

Date _____

Period _____

Acids And Bases Ws #6: More Fun Practice

	Solute	Name of solute	Molar Mass (g/ mole)	Molarity (mol/L)	grams solute Liter solution	[H ⁺]	[OH ⁻]	pH	pOH	acid or base
1.	HCl	Hydrochloric Acid	36.5	0.020	0.73	2.0×10^{-2}	5.0×10^{-13}	1.70	12.3	Acid
2.	NaOH			0.040						
3.	HClO ₄	Perchloric Acid	100.	1.25×10^{-3}	.125	1.25×10^{-3}	8×10^{-12}	2.90	11.1	Acid
4.	HNO ₃				0.060					
5.	RbOH	Rubidium hydroxide	102.5	3.75×10^{-2}	3.84	2.67×10^{-13}	3.75×10^{-2}	12.57	1.4^{-2}	Base
6.	H ₂ S			0.0030						
7.	HNO ₃	Nitric Acid	63.0	3×10^{-4}	.0189	3×10^{-4}	3.33×10^{-11}	3.52	10.48	Acid
8.	KOH						4.5×10^{-2}			
9.	Sr(OH) ₂	Strontium hydroxide	122	6.75×10^6	8.24×10^{-4}	7.41×10^{-10}	1.35×10^{-5}	9.13	4.87	Base