

Name \_\_\_\_\_ Date \_\_\_\_\_ Period \_\_\_\_\_



## Chemical Reactions Ws #2: Balancing Equations

- $2 \text{NaCl} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + 2 \text{HCl}$
- $\text{CuSO}_4 + 2 \text{KCl} \rightarrow \text{K}_2\text{SO}_4 + \text{CuCl}_2$
- $\text{BaCl}_2 + \text{Na}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + 2 \text{NaCl}$
- $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$
- $\text{AlCl}_3 + 3 \text{NH}_4\text{OH} \rightarrow 3 \text{NH}_4\text{Cl} + \text{Al(OH)}_3$
- $\text{AsCl}_3 + 3 \text{H}_2\text{O} \rightarrow 3 \text{HCl} + \text{As(OH)}_3$
- $\text{Cl}_2 + 2 \text{LiI} \rightarrow 2 \text{LiCl} + \text{I}_2$
- $\text{Sn} + 2 \text{KOH} \rightarrow \text{K}_2\text{SnO}_2 + \text{H}_2$
- $4 \text{FeS}_2 + 11 \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3 + 8 \text{SO}_2$
- $2 \text{MoO}_3 + 3 \text{Zn} + 3 \text{H}_2\text{SO}_4 \rightarrow \text{Mo}_2\text{O}_3 + 3 \text{ZnSO}_4 + 3 \text{H}_2\text{O}$
- $2 \text{PbO}_2(\text{s}) \rightarrow 2 \text{PbO}(\text{s}) + \text{O}_2(\text{g})$
- $4 \text{Al}(\text{s}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{Al}_2\text{O}_3(\text{s})$
- $2 \text{PbS}(\text{s}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{PbO}(\text{s}) + 2 \text{SO}_2(\text{g})$
- $\text{Zn(NO}_3)_2(\text{aq}) + \text{K}_2\text{S}(\text{aq}) \rightarrow \text{ZnS}(\text{s}) + 2 \text{KNO}_3(\text{aq})$
- $\text{Sn}(\text{s}) + 2 \text{HF}(\text{s}) \rightarrow \text{SnF}_2(\text{s}) + \text{H}_2(\text{g})$
- $\text{Cu}(\text{s}) + 2 \text{AgNO}_3(\text{aq}) \rightarrow \text{Cu(NO}_3)_2(\text{aq}) + 2 \text{Ag}(\text{s})$
- $2 \text{KOH}(\text{aq}) + \text{H}_2\text{SO}_4(\text{aq}) \rightarrow \text{K}_2\text{SO}_4(\text{aq}) + 2 \text{H}_2\text{O}(\text{l})$
- $4 \text{NH}_3(\text{g}) + 5 \text{O}_2(\text{g}) \rightarrow 4 \text{NO}(\text{g}) + 6 \text{H}_2\text{O}(\text{g})$

