



Chemical Reactions Ws #7: Writing Descriptions of Chem Rxns

Balance each of the following equations **and** write a description of each reaction using the correct compound names and physical states.



Solid Iron combines with oxygen gas to produce solid Iron III oxide



Solid Sodium combines with liquid bromine to produce solid sodium bromide



Solid Potassium combines with nitrogen gas to make solid potassium nitride



Solid Phosphorus combines with chlorine gas to produce liquid phosphorus chloride



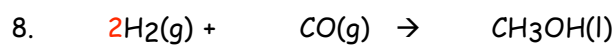
Solid aluminum oxide decomposes to yield solid aluminum and oxygen gas



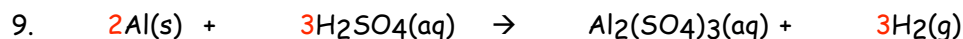
Solid Boron combines with chlorine gas to produce liquid boron chloride



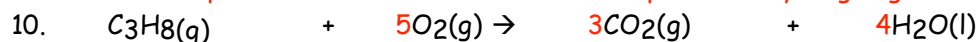
Solid sodium chlorate decomposes to produce solid sodium chloride and oxygen gas



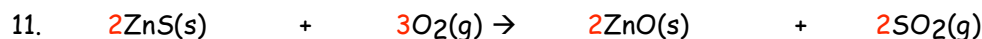
Hydrogen gas and carbon monoxide gas combine to produce liquid methanol



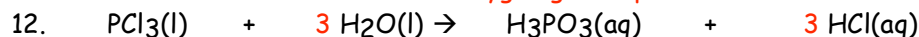
Solid Aluminum placed in a solution of sulfuric acid produces hydrogen gas & Aluminum sulfate solution



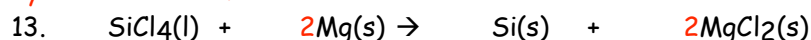
Propane combines with oxygen in a combustion reaction to produce carbon dioxide gas and liquid water



Solid Zinc II sulfide combines with oxygen gas to produces ZincII Oxide solide and sulfur dioxide gas.



Liquid Phosphorus chloride combines with water to produce an aqueous solution of phosphoric acid and hydrochloric acid.



Solid Magnesium is placed liquid silicon chloride and the Magnesium replaces the silicon in a single replacement reaction to yield solid silicon and solid magnesium chloride



Maganese IV Oxide solid in an aqueous solution of hydrochloric acid produces manganese II chloride, chlorine gas and liquid water.



Solutions of Lead II nitrate and potassium chromate combine in a double replacement reaction to produce solid lead II chromate and a solution of potassium nitrate.