

Name \_\_\_\_\_ Date \_\_\_\_\_ Period \_\_\_\_\_

### Chemical Reactions Ws #8: More Reaction and Balancing Practice

Write a **balanced** chemical equation for each of the following & determine the **type** of each reaction.

1. potassium + calcium fluoride →



1. Single replacement

2.  $Al(OH)_3 + H_2SO_4 \rightarrow Al_2(SO_4)_3 +$



2. Double replacement

3.  $C_2H_4 + O_2 \rightarrow$



3. Combustion

4. silver + sulfur → silver (I) sulfide



4. Single replacement

5.  $Al + Fe_2O_3 \rightarrow Fe + Al_2O_3$



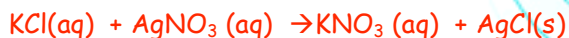
5. Single replacement

6. mercury(II) oxide → mercury + oxygen



6. Decomposition

7. potassium chloride + silver nitrate →



7. Double replacement

8.  $Ca + HCl \rightarrow$



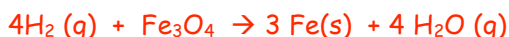
8. Single replacement

9.  $C_2H_5OH + O_2 \rightarrow$



9. Combustion

10. hydrogen +  $Fe_3O_4 \rightarrow$  iron + steam



10. Single replacement

11. carbon + oxygen → carbon monoxide



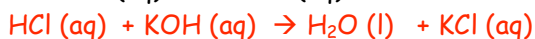
11. Combustion

12.  $Al_2(SO_4)_3 + Ca(OH)_2 \rightarrow$



12. Double replacement

13.  $HCl(aq) + KOH(aq) \rightarrow$



13. Double replacement/  
Neutralization

