

Name: _____ Date: _____ Period: _____

Naming WS #1: Ionic Compounds

	Ionic Formula	Name of Ionic Compound	Balanced Ion Pairs
1.	$Al_2 O_3$	Aluminum Oxide	$2 Al^{3+} + 3 O^{2-}$
2.		barium nitride	
3.	$Cs_2 O$	Cesium Oxide	$2 Cs^{1+} + O^{2-}$
4.		indium fluoride	
5.	MgO	Magnesium Oxide	$Mg^{2+} + O^{2-}$
6.		calcium oxide	
7.	$Rb_3 N$	Rubidium Nitride	$3 Rb^{+} + N^{3-}$
8.			$Ca^{2+} + 2Cl^{-}$
9.	Na₃N	sodium nitride	$3Na^{+} + N^{3-}$
10.	$Ca_3 N_2$		
11.	Li₂O	Lithium Oxide	$2Li^{+} + O^{2-}$
12.		Magnesium chloride	
13.	$Sr_2 Se$	Strontium Selenide	$2Sr^{+} + Se^{2-}$
14.			$3Na^{+} + P^{3-}$
15.	$Cs_2 S$	Cesium Sulfide	$2 Cs^{+} + S^{2-}$
16.			$Be^{2+} + 2Cl^{-}$
17.	K₂O	Potassium oxide	$2K^{+} + O^{2-}$
18.	$Al_2 S_3$		

19. Explain the difference between a cation and an anion. Explain how each is formed. Provide an
 Cation is a positive ion, it has lost an electron. Anion is a negative ion, it has gained an electron

20. In your own words, explain how to balance ion pairs. (Pretend you are teaching someone who has no background in atoms, ions, charges, losing/gaining electrons)