

Name: _____ Date: _____ Period: _____

Naming WS #3: Ionic Compounds with Transition Metals

Part A Write the formula for each ionic compound.

Part B Write the Name for each formula
Be sure to include the roman numeral.

	Name	Formula
1.	Tin (II) Hydroxide	Sn(OH)_2
2.	Iron (III) oxalate	
3.	Colbalt (II) bromide	CoBr_2
4.	Chromium (III) Chloride	
5.	Iron (II) Oxide	FeO
6.	Mercury (II) Chloride	
7.	Tin (IV) Sulfide	SnS_2
8.	Silver (I) Phosphide	
9.	Lead (IV) Iodide	PbI_4

	Formula	Name
1.	$\text{Fe(NO}_3)_3$	Iron III Nitrate
2.	Mn(OH)_2	
3.	$\text{Ti(NO}_3)_4$	Tin IV Nitrate
4.	PbS_2	
5.	CuCl	Copper I Chloride
6.	PbSO_4	
7.	ZnCl_2	Zinc II Chloride
8.	Hg_2O	
9.	Cr_2O_3	Chromium III Oxide

Part C: Fill in the following table. Polyatomic ions and transition metals are mixed up.

	Ionic Formula	Ionic Compound Name	Balanced Ion Pairs
1.	Na_2CrO_4	Sodium Chromate	$2\text{Na}^+ + \text{CrO}_4^{2-}$
2.	CuCl		
3.	NaNO_3	Sodium Nitrate	$\text{Na}^+ + \text{NO}_3^-$
4.	ZnSO_4		
5.	KCN	Potassium cyanide	$\text{K}^+ + \text{CN}^-$
6.	Al_2O_3		
7.	$\text{Cu(NO}_3)_2$	Copper(II) nitrate	$\text{Cu}^{2+} + 2\text{NO}_3^-$
8.	FePO_4		
9.	PbS	Lead(II) sulfide	$\text{Pb}^{2+} + \text{S}^{2-}$
10.	NaClO		
11.	Na_2CO_3	Sodium Carbonate	$2\text{Na}^+ + \text{CO}_3^{2-}$
12.	Na_2SeO_4		
13.	NH_4BrO_3	Ammonium Bromate	$\text{NH}_4^+ + \text{BrO}_3^-$
14.	$\text{Au(ClO}_4)_3$		
15.	$\text{Fe}_2(\text{SO}_4)_3$	Iron(III) Sulfate	$2\text{Fe}^{3+} + 3\text{SO}_4^{2-}$