

## Naming Practice Test

- The correct name for  $\text{LiCl}$  is
  - monolithium chloride
  - monolithium monochloride
  - lithium monochloride
  - lithium chloride
  - lithium(I) chloride
- The correct name for  $\text{FeO}$  is
  - iron(III) oxide
  - iron oxide
  - iron(I) oxide
  - iron monoxide
  - iron(II) oxide
- The charge on a potassium ion in its ionic compounds is
  - 2+
  - Various charges are possible.
  - 1+
- The correct name for the  $\text{Ca}^{2+}$  species is
  - calcium ion
  - calcium
  - monocalcium ion
  - calcium(II) ion
  - calcium(I) ion
- Titanium(IV) oxide has the formula
  - $\text{Ti}_4\text{O}_2$
  - $\text{TiO}_2$
  - $\text{TiO}_4$
  - $\text{Ti}_4\text{O}$
  - $\text{Ti(IV)O}$
- Which of the following formulas is *incorrect*?
  - $\text{Ba(OH)}_2$
  - $\text{CaCl}$
  - $\text{LiH}$
  - $\text{KMnO}_4$
  - $\text{K}_2\text{O}$
- The correct name for  $\text{P}_2\text{O}_5$  is
  - phosphorus pentoxide
  - phosphorus(V) oxide
  - diphosphorus pentoxide
  - diphosphorus oxide
  - phosphorus(II) oxide
- The correct formula for the carbonate ion is
  - $\text{CO}_4^{2-}$
  - $\text{CO}_3^-$
  - $\text{CO}_3^{2-}$
  - $\text{CO}_4^-$
  - $\text{CO}_3^{3-}$

## Naming Practice Test

9. The name for the  $\text{NO}_3^-$  ion is
- nitrogen ion
  - nitrate ion
  - nitrous ion
  - nitric ion
  - nitrite ion
10. The correct name for an aqueous solution of  $\text{H}_3\text{PO}_4$  is
- phosphate acid
  - phosphoric acid
  - phosphorous acid
  - hydrophosphorus acid
  - hydrophosphoric acid
11. The correct name for an aqueous solution of HCN is
- cyanate acid
  - cyanic acid
  - hydrocyanic acid
  - hydrocyanous acid
  - cyanous acid
12. The correct name for an aqueous solution of  $\text{H}_2\text{SO}_4$  is
- hydrosulfuric acid
  - hydrosulfurous acid
  - sulfurous acid
  - sulfuric acid
  - None of the above
13. The correct formula for ammonium sulfate is
- $(\text{NH}_4)_2\text{SO}_4$
  - $(\text{NH}_3)_2\text{SO}_3$
  - $\text{NH}_4\text{SO}_4$
  - $\text{NH}_4\text{SO}_3$
  - $(\text{NH}_4)_2\text{SO}_3$

Answer the following questions as a pair.

14. Which of the following is the correct formula for bromic acid?
- $\text{HBrO}_2(aq)$
  - $\text{HBrO}_4(aq)$
  - $\text{HBr}(aq)$
  - $\text{HBrO}_3(aq)$
  - $\text{HBrO}(aq)$
15. The name for  $\text{NH}_4\text{Br}$  is
- ammonium bromide
  - nitrogen hydrogen bromide
  - ammonium(I) bromide
  - nitrogen tetrahydrogen bromide
  - None of the above

## Naming Practice Test

16. The name for  $\text{Ba}(\text{NO}_3)_2$  is
- A. barium(II) nitrate
  - B. barium nitrate
  - C. barium(I) nitrate
  - D. barium dinitrate
  - E. barium nitrite
17. The correct name for  $\text{Cu}_2\text{O}$  is
- A. copper oxide
  - B. dicopper oxide
  - C. dicopper monoxide
  - D. copper(I) oxide
  - E. copper(II) oxide
18. Write the correct formula for ammonium dichromate.
19. Write the correct formula for beryllium hydride.
20. What is the formula for sulfur hexafluoride?
- A.  $\text{SF}_4$
  - B.  $\text{S}_3\text{F}$
  - C.  $\text{S}_6\text{F}$
  - D.  $\text{SF}_5$
  - E.  $\text{SF}_6$
21. What is the formula for nickel(II) carbonate?
- A.  $\text{NiC}$
  - B.  $\text{Ni}(\text{CO}_3)_2$
  - C.  $\text{NiCO}_3$
  - D.  $\text{Ni}_2\text{CO}_3$
  - E.  $\text{Ni}_2(\text{CO}_3)_3$
22. Write the correct formula for iron(III) sulfide.
23. Write the correct formula for lead(IV) oxide.
24. Give the name for  $\text{SiO}_2$ .
25. Give the name for  $\text{CO}_2$ .

**Answer Sheet**

1. D. lithium chloride
2. E. iron(II) oxide
3. C. 1+
4. A. calcium ion
5. B.  $\text{TiO}_2$
6. B.  $\text{CaCl}$
7. C. diphosphorus pentoxide
8. C.  $\text{CO}_3^{2-}$
9. B. nitrate ion
10. B. phosphoric acid
11. C. hydrocyanic acid
12. D. sulfuric acid
13. A.  $(\text{NH}_4)_2\text{SO}_4$
14. D.  $\text{HBrO}_3(\text{aq})$
15. A. ammonium bromide
16. B. barium nitrate
17. D. copper(I) oxide
18.  $(\text{NH}_4)_2\text{Cr}_2\text{O}_7$
19.  $\text{BeH}_2$
20. E.  $\text{SF}_6$
21. C.  $\text{NiCO}_3$
22.  $\text{Fe}_2\text{S}_3$
23.  $\text{PbO}_2$
24. silicon dioxide
25. carbon dioxide

**Standards Summary**

NSES B.2.3

Bonds between atoms are created when electrons are paired up by being transferred or shared. A substance composed of a single kind of atom is called an element. The atoms may be bonded together into molecules or crystalline solids. A compound is formed when two or more kinds of atoms bind together chemically.