

Name _____ Date _____ Period _____



Chemical Reactions Ws #2: Balancing Equations

- $\text{NaCl} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{HCl}$
- $\text{CuSO}_4 + \text{KCl} \rightarrow \text{K}_2\text{SO}_4 + \text{CuCl}_2$
- $\text{BaCl}_2 + \text{Na}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + \text{NaCl}$
- $\text{KClO}_3 \rightarrow \text{KCl} + \text{O}_2$
- $\text{AlCl}_3 + \text{NH}_4\text{OH} \rightarrow \text{NH}_4\text{Cl} + \text{Al}(\text{OH})_3$
- $\text{AsCl}_3 + \text{H}_2\text{O} \rightarrow \text{HCl} + \text{As}(\text{OH})_3$
- $\text{Cl}_2 + \text{LiI} \rightarrow \text{LiCl} + \text{I}_2$
- $\text{Sn} + \text{KOH} \rightarrow \text{K}_2\text{SnO}_2 + \text{H}_2$
- $\text{FeS}_2 + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3 + \text{SO}_2$
- $\text{MoO}_3 + \text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{Mo}_2\text{O}_3 + \text{ZnSO}_4 + \text{H}_2\text{O}$
- $\text{PbO}_2(\text{s}) \rightarrow \text{PbO}(\text{s}) + \text{O}_2(\text{g})$
- $\text{Al}(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{Al}_2\text{O}_3(\text{s})$
- $\text{PbS}(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{PbO}(\text{s}) + \text{SO}_2(\text{g})$
- $\text{Zn}(\text{NO}_3)_2(\text{aq}) + \text{K}_2\text{S}(\text{aq}) \rightarrow \text{ZnS}(\text{s}) + \text{KNO}_3(\text{aq})$
- $\text{Sn}(\text{s}) + \text{HF}(\text{s}) \rightarrow \text{SnF}_2(\text{s}) + \text{H}_2(\text{g})$
- $\text{Cu}(\text{s}) + \text{AgNO}_3(\text{aq}) \rightarrow \text{Cu}(\text{NO}_3)_2(\text{aq}) + \text{Ag}(\text{s})$
- $\text{KOH}(\text{aq}) + \text{H}_2\text{SO}_4(\text{aq}) \rightarrow \text{K}_2\text{SO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l})$
- $\text{NH}_3(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{NO}(\text{g}) + \text{H}_2\text{O}(\text{g})$