

Name \_\_\_\_\_ Date \_\_\_\_\_ Period \_\_\_\_\_

### Mole Unit WS #6: More Mole Conversion Practice

1. Define "mole" in terms of number and mass.
2. How many moles are present in 34.0 grams of  $\text{Cu}(\text{OH})_2$ ?
3. How many moles are present in  $2.45 \times 10^{23}$  molecules of  $\text{CH}_4$ ?
4. What is the mass of  $3.4 \times 10^{24}$  molecules of  $\text{NH}_3$ ?
5. What is the mass of 4.20 moles of  $\text{Ca}(\text{NO}_3)_2$ ?
6. What is the molar mass of  $\text{MgO}$ ?
7. How many moles are represented by 16.0 g of ethanol,  $\text{C}_2\text{H}_5\text{OH}$ ?
8. How many grams of glucose are in  $6.63 \times 10^{23}$  molecules of glucose,  $\text{C}_6\text{H}_{12}\text{O}_6$ ?
9. How many moles of methanol,  $\text{CH}_3\text{OH}$ , are there in  $6.53 \times 10^{23}$  molecules of methanol?
10. How many moles of  $\text{NaCl}$  are in 16.0 grams of  $\text{NaCl}$ ?
11. How many molecules of potassium hydroxide,  $\text{KOH}$  are in 40.6 g?
12. How many molecules are there in 450 grams of  $\text{Na}_2\text{SO}_4$ ?