

## Naming WS #2: Ionic Compounds with Polyatomic Ions

Part A - Complete the following table.

	Ionic Compound:	Ion Pairs:	Name of Ionic Compound:
Ex.	$\text{AlCl}_3$	$\text{Al}^{3+} + 3\text{Cl}^-$	Aluminum chloride
1.			ammonium perchlorate
2.		$\text{Li}^+ + \text{NO}_2^-$	
3.	$\text{Ga}_2(\text{SO}_3)_3$		
4.		$2\text{Li}^+ + \text{CO}_3^{2-}$	
5.			calcium chlorate
6.	$\text{Na}_2\text{SiO}_3$		
7.			barium hydroxide
8.		$\text{Be}^{2+} + \text{C}_2\text{O}_4^{2-}$	
9.			lithium cyanide
10.	$(\text{NH}_4)_3\text{PO}_4$		

Part B- Indicate the names of the ions used to form the following ionic compounds and give the ionic compound name of each.

	Formula	Cation name		Anion Name		Compound name
Ex	$\text{NH}_4\text{NO}_3$	Ammonium ion	+	Nitrate ion	=	Ammonium Nitrate
11.	$\text{Na}_2\text{SO}_4$		+		=	
12.	$\text{K}_3\text{PO}_4$		+		=	
13.	$\text{Ba}(\text{NO}_3)_2$		+		=	
14.	$\text{Al}(\text{OH})_3$		+		=	
15.	$\text{Na}_2\text{SO}_3$		+		=	
16.	$\text{NaHCO}_3$		+		=	
17.	$\text{CsClO}_4$		+		=	